

# Electrochemistry

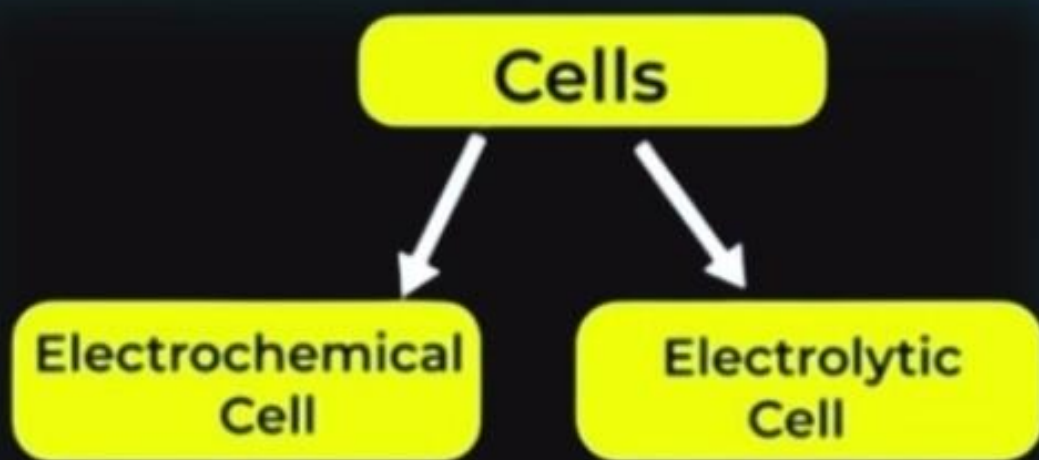
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*Electrochemistry is the subdiscipline of Chemistry that deals with the study of the relationship between electrical energy and chemical changes.*

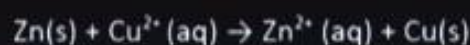
*chemical energy ---> electric energy*  
*electric energy ---> chemical energy*



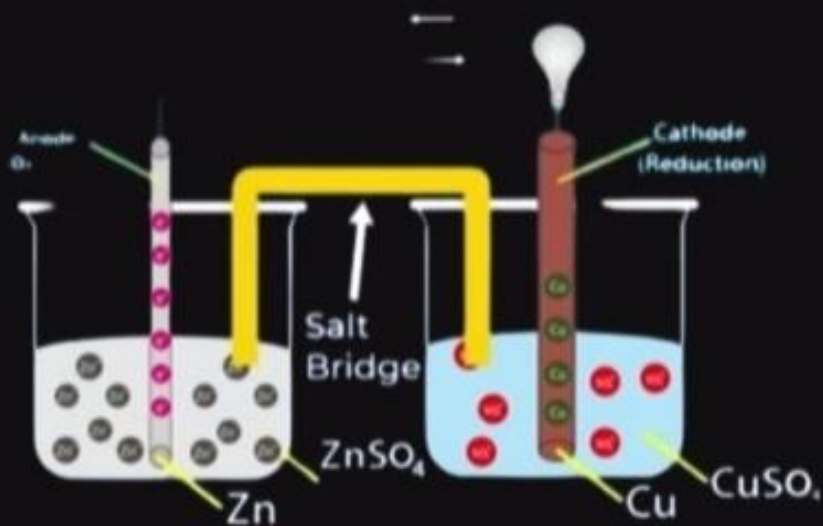
Chemical energy to  
Electrical energy  
Also called **Galvanic**  
Or  
**Voltaic Cell**

# Electrochemical cell

👉 *An electrochemical cell is a device that can generate electrical energy from the chemical reactions occurring in it, or use the electrical energy .*



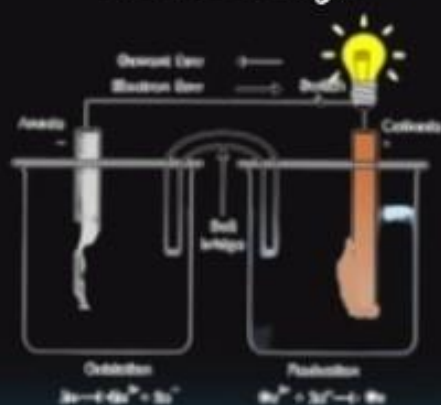
# Daniell Cell





# Galvanic Cell

*A galvanic cell is generally is represented by putting a vertical line between metal and electrolyte solution and putting a double vertical line between two electrolyte connected by a salt bridge*



Galvani



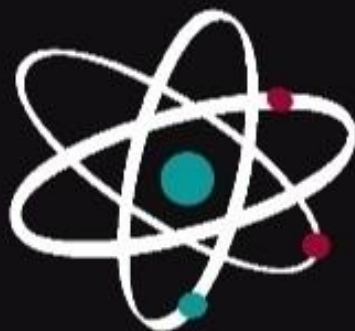
## cell potential / EMF

\* potential difference between two electrodes of a galvanic cell.

Measuring volt

\*  $e(\text{cell})$  called the cell electromotive force then no current is drawn through the cell .

**Thank you**



**from : chemistry department**

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